

REMARKS

Reconsideration of this application, as amended, is respectfully requested.

Claims 32-42 were rejected under 35 U.S.C. §103(a) for allegedly being obvious over Dolan in view of Schaedel. Applicants respectfully traverse.

Dolan discloses an anodizing method for light metals whereby solutions of different compositions may be used (embodiments A to G, cols. 4-9). In the list of substance classes of col. 2, middle, a), b) and i) are respectively phosphorus oxysalts, silicon oxysalts and alkali metal hydroxides. Surfactants, alcohols nor alkaline hydrolyzed silanes are not mentioned.

Schaedel describes an anodizing process for aluminum parts, whereby the aluminum oxide coating is formed during the anodizing process by spraying with and not by dipping the part into the anodizing solution. The spray causes a foam surrounding the part. Therefore, it is possible to use very long parts to be anodized without the need to construct bath containers for which dimensions (col. 2, lines 22-24). This solution contains a surfactant and optionally a carboxylic acid and either a triethanolamine or an alcohol or both. The alcohols may be heavy complex alcohols like glycerine, mannitol and sorbitol (claim 23).

The alkaline hydrolyzed silanes of the present application cannot be mixed with the silicon oxysalts of Dolan, as these are silicates.

The alcohols of the present application are alcohols having at least one alkaline radical group, whereas Schaedel mentions heavy complex alcohols like glycerine, mannitol and sorbitol in col. 13, lines 16/17. The alcohols of Schaedel have a high viscosity to easily generate foam together with a surfactant and should have a pH of about 7 to 8, whereas the alcohols of the present application should have a very low viscosity and should be alkaline, generally about pH 11.

The surfactant added to the solution of the present application should not create foam because it is intended to work with a bath where the articles to be coated are dipped into it, and any foam would disturb this operation as the foam may negatively influence the electrical system. The foam may cause discharges and short circuits which negatively influence the quality of the coating to be generated and which may even destroy the equipment used in the process.

In view of the foregoing, this rejection should be withdrawn.

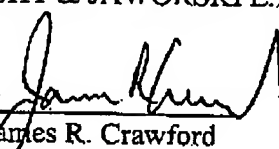
Claims 32-42 were rejected on the ground of obviousness-type double patenting over U.S. Patent No. 6,875,334 in view of Kinase. Applicants respectfully traverse.

Triethanolamine is not an accelerator for generating the phosphate containing coating like hydroxylamine. Triethanolamine may be one of the alcohols which is used in the present application for the stabilization of the gel for the polymeric layer. Hydroxylamine would not work as a stabilizer in the solution of the process of the present application. Thus, the Examiner has not considered the use of the respective components, and it would not be obvious to substitute one for the other.

Thus, the obviousness-type double patenting rejection should be withdrawn.

Any necessary fees may be charged to deposit account no. 50-0624.

Respectfully submitted
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